Gas-phase reaction of fullerene monocations with 2,3-benzofluorene indicates the importance of charge exchanges

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ABSTRACT

Fullerene and polycyclic aromatic hydrocarbon (PAH) molecules, as well as their cations and clusters, are of great interest in astrochemistry. In this work, the ion-molecule collision reaction between fullerene (e.g. a C_{60}/56/58 and C_{60} System or a C_{64/66/68} and C_{70}) monocations and neutral PAHs (e.g. 2,3-benzofluorene, C_{17}H_{12}) is studied in the gas phase to determine the importance of charge exchanges and to illustrate the competition between charge transfer and molecular adduct formation channels. The experimental results show that the charge transfer channel is the dominant channel (i.e. charge exchange) in the reaction between fullerene (C_{60} and C_{64}) monocations and 2,3-benzofluorene, while the molecular adduct formation channels are the dominant channels in the reaction between fullerene (C_{54/56/58} and C_{64/66/68}) monocations and 2,3-benzofluorene. The observed reaction behaviours are investigated with quantum calculations, and the CH\textsubscript{2} unit binding effect of 2,3-benzofluorene is determined to be the main reason for the results. Our findings on the ion-molecule collision reaction between fullerene monocations and 2,3-benzofluorene provide a good model for understanding the physical-chemical processes of the charge transfer channel and the cluster adduct formation channels. Neutral fullerenes (C_{60} and C_{64}) increase the abundance of their monocations through collision reactions with coexisting neutral molecules in the interstellar medium.

Key words. astrochemistry — methods: laboratory — ultraviolet: ISM — ISM: molecules — molecular processes

1. Introduction

The buckminsterfullerene (C_{60}) molecule was first ‘discovered’ in the laboratory in 1985 (Kroto et al. 1985). In 2010, the existence of buckminsterfullerene (C_{60}) in space was discovered through the infrared (IR) spectra of circumstellar and interstellar sources (Cami et al. 2010; Sellgren et al. 2010). Importantly, for the first time, several far-red diffuse interstellar bands were linked to the electronic transitions of buckminsterfullerene monocations (C_{60}^+) (Campbell et al. 2015; Walker et al. 2015; Cordiner et al. 2017). Hence, understanding the processes that regulate the origin and evolution of these fullerene species and their relationships with the organic inventory in space has become a focus of astrochemistry research (Tielens 2013, and references therein).

Fullerenes are exceptionally stable carbon cages that are difficult to destroy or modify (Handschuh et al. 1995; Candian et al. 2019). However, minor modifications that preserve the cage size and composition are much easier to accomplish, such as ionization, isomerization, and exohedral atom addition, as well as other chemical reactions, including association with PAH molecules (Jäger et al. 2009, 2011; Zhen et al. 2019b). Therefore, when the total abundance of a fullerene species, such as C_{60}, in an interstellar region is considered, its different ionization or chemical states should be considered instead of its abundance in a given state in space (Omont 2016; Cordiner et al. 2019, and references therein).

As C_{60} is an important component of the interstellar organic inventory, various studies have been performed on the formation mechanism of C_{60}, and the formation of fullerenes has been shown to be chemically linked to large PAHs in the interstellar medium (ISM) (Berné & Tielens 2012; Zhen et al. 2014; Andrews et al. 2015). Given the presence of C_{60} and PAHs in photo-dissociation regions (PDRs), such as NGC 7023 (Sellgren et al. 2010; Berné & Tielens 2012), studies on the chemical evolution network of C_{60} and PAH species have piqued the interest of researchers (for example, see Bohme (2016); Candian et al. (2019); Zhen et al. (2019a,b)).

The adduction reactions of large molecular clusters and the charge transfer reactions of fullerene and PAHs have been extensively studied by Bohme and colleagues using the selected ion flow tube technique. The reactions of singly, doubly, and triply charged C_{60} with a wide range of collision partners have been studied (Bohme 1992; Petrie & Bohme 1993; Bohme 2016). The adduction reactions compete with the charge transfer channels, and doubly and triply charged fullerene (C_{60}) cations exhibit considerably different behaviours relative to singly charged fullerene (C_{60}) cations. The main reason for this is that charge transfer is endothermic for reactions with singly charged cations and exothermic for reactions with doubly and
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2. Experimental results

The experiments were performed on an apparatus equipped with a quadrupole ion trap and reflection time-of-flight mass spectrometer, and the experimental details are provided in the appendix A (Zhen et al. 2019a,b). During the experiments, the high energy of the impacting electrons (∼82 eV) causes fullerene (C_{60} and C_{70}) monocations to form, and the fragmentation of the original fullerene cations through C_2 losses leads to the formation of smaller fullerenes (C_{54/56/58} or C_{64/66/68}) monocations (Lifshitz 2000; Zhen et al. 2014). The resulting mass spectra are presented in Figures 1-2.

The mass spectra of the fullerene (C_{54/56/58} and C_{60}) monocation collision reactions with (red line) and without (blue line) neutral 2,3-benzofluorene are shown in Figure 1(A). Clearly, during the reaction with 2,3-benzofluorene, a strong mass peak associated with the 2,3-benzofluorene cations C_{17}H_{12}^+ (m/z=216) is observed, as well as a series of weak mass peaks associated with the fullerene/2,3-benzofluorene cluster cations. In Figure 1(B), we label the newly formed fullerene (C_{54/56/58}, C_{56}, C_{58}, and C_{60})/2,3-benzofluorene cluster cations as [(C_{17}H_{12})C_{54}]^+ (m/z=864), [(C_{17}H_{12})C_{56}]^+ (m/z=888), [(C_{17}H_{12})C_{58}]^+ (m/z=912), and [(C_{17}H_{12})C_{60}]^+ (m/z=936), respectively.
Fig. 2. Mass spectrum of the products resulting from the fullerene (C\textsubscript{64/66/68} and C\textsubscript{70}) monocation collision reaction with and without neutral 2,3-benzofluorene. Upper panel (A): Mass spectrum of fullerene (C\textsubscript{64/66/68} and C\textsubscript{70}) monocations with and without 2,3-benzofluorene. Panel (B): Presence of [C\textsubscript{64/66/68/70}]\textsuperscript{+} and the newly formed [(C\textsubscript{17}H\textsubscript{12})C\textsubscript{64/66/68/70}]\textsuperscript{+} cluster cations.

Importantly, the intensity of the species I(C\textsubscript{56}\textsuperscript{+}, without 2,3-benzofluorene) ≃ I(C\textsubscript{58}\textsuperscript{+}, with 2,3-benzofluorene) + I([C\textsubscript{17}H\textsubscript{12}C\textsubscript{58}]\textsuperscript{+}), and I(C\textsubscript{56}\textsuperscript{+}, without 2,3-benzofluorene) ≃ I(C\textsubscript{56}\textsuperscript{+}, with 2,3-benzofluorene) + I([C\textsubscript{17}H\textsubscript{12}C\textsubscript{56}]\textsuperscript{+}). This result suggests that for fullerene (C\textsubscript{56/58}) monocations, the molecular cluster formation channel is the dominant channel. Furthermore, the intensity of the species I(C\textsubscript{60}\textsuperscript{+}, without 2,3-benzofluorene) ≃ I(C\textsubscript{60}\textsuperscript{+}, with 2,3-benzofluorene) + I([C\textsubscript{17}H\textsubscript{12}C\textsubscript{60}]\textsuperscript{+}) + C\textsubscript{17}H\textsubscript{12}\textsuperscript{+}, and the intensity ratio of C\textsubscript{17}H\textsubscript{12}\textsuperscript{+}:[(C\textsubscript{17}H\textsubscript{12})C\textsubscript{60}]\textsuperscript{+} ≃ 100: 1, which suggests that, for fullerene (C\textsubscript{60}) monocations, the charge transfer channel is the dominant channel. The intensity of the newly formed C\textsubscript{17}H\textsubscript{12}\textsuperscript{+} is similar to the intensity of I(C\textsubscript{60}\textsuperscript{+}, without 2,3-benzofluorene) - I(C\textsubscript{60}\textsuperscript{+}, with 2,3-benzofluorene) as a result of the ion-molecular charge transfer channel. We discuss this result later in combination with the theoretical results.

Based on the above results, for the gas-phase reaction of fullerene (C\textsubscript{34/56/58} and C\textsubscript{60}) monocations with neutral 2,3-benzofluorene, the reaction pathways are as follows. The fullerene (C\textsubscript{60}) monocation reaction with neutral 2,3-benzofluorene is

\[
[C\textsubscript{60}]\textsuperscript{+} + C\textsubscript{17}H\textsubscript{12} \rightarrow [C\textsubscript{17}H\textsubscript{12}C\textsubscript{60}]\textsuperscript{+} \sim 1\% \\
[C\textsubscript{60}]\textsuperscript{+} + C\textsubscript{17}H\textsubscript{12} \rightarrow [C\textsubscript{17}H\textsubscript{12}]\textsuperscript{+} + C\textsubscript{60} \sim 99\%.
\]

The fullerene (C\textsubscript{54/56/58}) monocation reaction with neutral 2,3-benzofluorene is

\[
[C\textsubscript{54/56/58}]\textsuperscript{+} + C\textsubscript{17}H\textsubscript{12} \rightarrow [C\textsubscript{17}H\textsubscript{12}C\textsubscript{54/56/58}]\textsuperscript{+} \sim 100\%.
\]

Similar to the reaction behaviour of fullerene (C\textsubscript{60} and C\textsubscript{54/56/58}) monocation and 2,3-benzofluorene systems, the mass spectra of the fullerene (C\textsubscript{64/66/68} and C\textsubscript{70}) monocation collision reaction with (red line) and without (blue line) neutral 2,3-benzofluorene are shown in Figure 2(A). Clearly, when 2,3-benzofluorene is present, a strong mass peak associated with the 2,3-benzofluorene cluster cations [C\textsubscript{17}H\textsubscript{12}]\textsuperscript{+} (m/z = 216) is observed, as well as a series of weak mass peaks associated with the fullerene/2,3-benzofluorene cluster cations. As shown in Figure 2(B), we label the newly formed fullerene (C\textsubscript{64}, C\textsubscript{66}, C\textsubscript{68}, and C\textsubscript{70})/2,3-benzofluorene cluster cations as [(C\textsubscript{17}H\textsubscript{12}C\textsubscript{66}]\textsuperscript{+} (m/z = 984), [(C\textsubscript{17}H\textsubscript{12}C\textsubscript{68}]\textsuperscript{+} (m/z = 1008), [(C\textsubscript{17}H\textsubscript{12}C\textsubscript{68}]\textsuperscript{+} (m/z = 1032), and [(C\textsubscript{17}H\textsubscript{12}C\textsubscript{70}]\textsuperscript{+} (m/z = 1056), respectively.

Importantly, the intensity of the species I(C\textsubscript{68}\textsuperscript{+}, without 2,3-benzofluorene) ≃ I(C\textsubscript{68}\textsuperscript{+}, with 2,3-benzofluorene) + I([(C\textsubscript{17}H\textsubscript{12}C\textsubscript{68}])\textsuperscript{+}) and I(C\textsubscript{66}\textsuperscript{+}, without 2,3-benzofluorene) ≃ I(C\textsubscript{66}\textsuperscript{+}, with 2,3-benzofluorene) + I([(C\textsubscript{17}H\textsubscript{12}C\textsubscript{66})]). This result suggests that for fullerene (C\textsubscript{66/68}) monocations, the molecular
cluster formation channel is the dominant channel. Furthermore, 
the intensity of the species $I(C_{70}^+, \text{with } 2,3\text{-benzofluorene}) \approx I(C_{70}^+, \text{with } 2,3\text{-benzofluorene}) + I((C_{17}H_{12})C_{70}^+ \text{ or } C_{17}H_{12}^+, \text{with } 2,3\text{-benzofluorene})$, and the intensity ratio of $C_{17}H_{12}^+((C_{17}H_{12})C_{70}^-) \approx 99:1$, which suggests that, for fullerene ($C_{70}$) monocations, the charge transfer channel is the dominant channel. The intensity of the newly formed $C_{17}H_{12}^+$ is similar to the intensity of $I(C_{70}^+, \text{without } 2,3\text{-benzofluorene}) - I(C_{70}^+, \text{with } 2,3\text{-benzofluorene})$ due to the ion-molecular charge transfer channel. We discuss this result later in combination with the theoretical results.

Based on the above results, for the gas-phase reaction of fullerene ($C_{64/66/68}$ and $C_{58}$) monocations with neutral 2,3-benzofluorene, we propose the following reaction pathways. The fullerene ($C_{70}$) monocation reaction with neutral 2,3-benzofluorene is

$$[C_{70}^+] + C_{17}H_{12} \rightarrow [C_{17}H_{12}C_{70}^+] \sim 1\%$$

$$[C_{70}^+] + C_{17}H_{12} \rightarrow [C_{17}H_{12}^+] + C_{70} \sim 99\%.$$ 

The fullerene ($C_{64/66/68}$) monocation reaction with neutral 2,3-benzofluorene is

$$[C_{64/66/68}^+] + C_{17}H_{12} \rightarrow [C_{17}H_{12}C_{64/66/68}^+] \sim 100\%.$$ 

In addition, in Figures 1 and 2, we observe additional mass features when 2,3-benzofluorene is not present. We suspect that these mass peaks, which occur at approximately m/z=200 in Figures 1 and 2, form either due to previous similar experiments or as a side product due to contamination in the vacuum chamber. The mass peaks at approximately m/z=360 in Figure 1 are the dication species of $C_{58/60}$; the mass peaks at approximately m/z=420 in Figure 2 are the dication species of $C_{68/70}$; and the mass peaks at approximately m/z=600 in Figure 2 are the cation species of previous experiments (dicoronylene cations, $C_{68}H_{20}^+$). We do not believe that these species affect the interpretation of our measurements due to their relatively small magnitudes.

3. Theoretical calculation results

To understand the details of the reaction process of the fullerene monocations with neutral 2,3-benzofluorene, we use $C_{60}^+ + 2,3\text{-benzofluorene}$ and $C_{58}^+ (7\text{ C-ring or 6 C-ring}) + 2,3\text{-benzofluorene}$ as examples and study their theoretical adduct reactions and charge transfer processes. We follow the minimum energy pathway. At each step, from the van der Waals cluster to the covalently bonded clusters, the energy and optimized structures are calculated. The details of the theoretical calculation are provided in the appendix A.

The energies and optimized structures for the reactant, transition states (TS, or TS1 and TS2), intermediate states, and product of the adduct reaction pathway between $C_{60}^+$ and 2,3-benzofluorene, $C_{58}^+$ (7 C-ring/6 C-ring) and 2,3-benzofluorene are shown in Figures 3-5, as well as the charge transfer channels. At the beginning of the calculations, we consider the reaction process between the fullerene cations and the 2,3-benzofluorene clusters in two types of channels. The I channel (the blue line) is the charge transfer channel, where charge exchange occurs during the collision reaction between the fullerene cations and 2,3-benzofluorene. The II channel (the red line) is the adduct formation channel, where clusters (van der Waals clusters or covalently bonded clusters) form during the collision reaction between the fullerene cations and 2,3-benzofluorene. These two channels compete throughout the reaction process, experiencing relatively similar intermediate processes or sharing the same initial precursors (fullerene cation + 2,3-benzofluorene, which is defined as 0 eV).

3.1. The optimized geometric structure of the fullerene ($C_{58}$) monocations and neutral 2,3-benzofluorene

We assume that the fullerene ($C_{58}$) monocations do not have a carbon skeleton rearrangement (except for the $C_2$ loss at a local position) during the electron impact ionization and fragmenta-

![Fig. 3. Reactants, transition states, intermediate states, products, and corresponding energies of the reaction pathways of $C_{60}^+$ with 2,3-benzofluorene: the I channel (the blue line) is the charge transfer channel, and the II channel (the red line) is the adduct formation channel. These two channels originate from the same starting point, which is defined as 0 eV.](image-url)
Zhen et al. (2019b). Following calculations (Lee & Han 2004; Candian et al. 2019), fullerene (C$_{60}$) is an open-shell doublet ground state with a spin multiplicity of 2 in the electronic ground state.

These two channels originate from the same starting point, which is defined as 0 eV.

**Fig. 4.** Reactants, transition states, intermediate states, products, and corresponding energies of the reaction pathways of C$_{60}^+$ (6 C-ring) with 2,3-benzofluorene: the I channel (the blue line) is the charge transfer channel, and the II channel (the red line) is the adduct formation channel. Two transition states are considered (Zhen et al. 2019b). The first pathway is in the '6 C-ring', and the other pathway is in the '7 C-ring'.

3.2. Reaction process of C$_{60}^+$ with C$_{17}$H$_{12}$

As shown in Figure 3, at the beginning of the addition process (red line), C$_{60}^+$ and 2,3-benzofluorene form a van der Waals molecular complex ([C$_{60}^+$-(C$_{17}$H$_{12}$)]) with an exothermic energy of approximately 1.00 eV. For the covalently bonded product, in the 2,3-benzofluorene ‘landing’ during the C$_{60}^+$ process, two carbon atoms from C$_{60}$ are removed from the cage surface, and the 2,3-benzofluorene structure is modified to allow the 9, 10 C-atoms to bond to the fullerene C-atoms (Sato et al. 2013; Zhen et al. 2019b). During the adduct reaction process, charge transfer (blue line) easily occurs between the C$_{60}$ monocation and neutral 2,3-benzofluorene. As presented in Figure 3, the charge exchange reaction of C$_{60}^+$ with 2,3-benzofluorene is exothermic with an energy of 0.27 eV, which is thermodynamically favourable.

3.3. Reaction of C$_{58}^+$ (6 C-ring) with C$_{17}$H$_{12}$

For the reaction of C$_{58}$ monocations with neutral 2,3-benzofluorene, due to the structure of C$_{58}^+$, two typical reaction pathways are considered (Zhen et al. 2019b). The first pathway is in the ‘6 C-ring’, and the other pathway is in the ‘7 C-ring’. The calculation results are presented in Figures 4 and 5, respectively.

The theoretical results of the reaction pathway between C$_{58}^+$ (6 C-ring) and 2,3-benzofluorene are presented in Figure 4. At the beginning of the addition process (red line), C$_{58}^+$ (6 C-ring) and 2,3-benzofluorene form a van der Waals molecular complex ([C$_{58}^+$ (6 C-ring)-(C$_{17}$H$_{12}$)]) with an exothermic energy of approximately 0.78 eV. For the covalently bonded product, in the 2,3-benzofluorene ‘landing’ in the C$_{58}^+$ process, two carbon atoms from C$_{58}$ are removed from the cage surface, and the 2,3-benzofluorene structure is modified to allow the 9, 10 C-atoms to bond to the C-atoms from C$_{58}^+$ (6 C-ring) (Sato et al. 2013; Zhen et al. 2019b). Due to the asymmetric structure of 2,3-benzofluorene, two transition states are obtained based on...
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Fig. 5. Reactants, transition states, intermediate states, products, and corresponding energies of the reaction pathways of C$_{58}$+ (7 C-ring) with 2,3-benzofluorene: the I channel (the blue line) is the charge transfer channel, and the II channel (the red line) is the adduct formation channel. These two channels originate from the same starting point, which is defined as 0 eV.

3.4. Reaction of C$_{58}$+ (7 C-ring) with C$_{17}$H$_{12}$

The theoretical results of the reaction pathway between C$_{58}$+ (7 C-ring) and 2,3-benzofluorene are presented in Figure 5. As shown in Figure 5, initially (red line) C$_{58}$+ (7 C-ring) and 2,3-benzofluorene form a covalently bonded molecule. After that, the final product (product, [C$_{58}$-(C$_{17}$H$_{12}$)]$^+$, -0.01 eV) is formed. However, considering the energy loss of van der Waals molecules, such as IR emissions and collisions, no covalently bonded molecules form during the transition (TS, [C$_{58}$-(a-C$_{17}$H$_{12}$)]$^+$ and TS, [C$_{58}$-(b-C$_{17}$H$_{12}$)]$^+$).

During the adduct reaction process, charge transfer (blue line) cannot occur between the C$_{58}$ cation and 2,3-benzofluorene. As presented in Figure 4, the charge exchange reaction of C$_{58}$+ with 2,3-benzofluorene is endothermic with an energy of 0.32 eV, which is thermodynamically unfavourable.

the CH$_2$ unit of 2,3-benzofluorene, as presented in Figures 4(A) and 4(B).

Figure 4(A) shows that the CH unit first bonds to the fullerene cage. Then, the CH unit goes through a transition state ([C$_{58}$-(a-C$_{17}$H$_{12}$)]$^+$, TS, +0.22 eV) and passes the first activation barrier (1.00 eV). Next, the final product (product, [C$_{58}$-(C$_{17}$H$_{12}$)]$^+$, -0.01 eV) is formed. As presented in Figure 4(B), the CH unit (the same side as the CH$_2$ unit) first bonds to the fullerene cage, then enters a transition state ([C$_{58}$-(b-C$_{17}$H$_{12}$)]$^+$, TS, +0.33 eV) and passes the first activation barrier (1.11 eV). After that, the final product (product, [C$_{58}$-(b-C$_{17}$H$_{12}$)]$^+$, -0.01 eV) is formed. However, considering the energy loss of van der Waals molecules, such as IR emissions and collisions, no covalently bonded molecules form during the transition (TS, [C$_{58}$-(a-C$_{17}$H$_{12}$)]$^+$ and TS, [C$_{58}$-(b-C$_{17}$H$_{12}$)]$^+$).

During the adduct reaction process, charge transfer (blue line) cannot occur between the C$_{58}$ cation and 2,3-benzofluorene. As presented in Figure 4, the charge exchange reaction of C$_{58}$+ with 2,3-benzofluorene is endothermic with an energy of 0.32 eV, which is thermodynamically unfavourable.
benzofluorene form a van der Waals molecular complex \([\text{C}_{58}(\text{7 C-ring})-(\text{a-C}_{17}\text{H}_{12})]^+\), with an exothermic energy of approximately 0.93 eV. For the covalently bonded product, in the 2,3-benzofluorene ‘landing’ during the \(\text{C}_{58}^+\) process, two carbon atoms from \(\text{C}_{58}\) are removed from the cage surface, and the 2,3-benzofluorene structure is modified to allow the 9, 10 C-atoms to bond to the C-atoms from \(\text{C}_{58}^+\) (Sato et al. 2013; Zhen et al. 2019b). Due to the asymmetric structure of 2,3-benzofluorene, multiple transition pathways are obtained based on the effect of the CH\(_2\) unit of 2,3-benzofluorene, as presented in Figures 5(A) and 5(B).

As presented in Figure 5(A), the CH unit first bonds to the fullerene cage (7 C-ring) first; then, the CH unit enters an intermediary state (Inter, \([\text{C}_{58}(\text{7 C-ring})-(\text{a-C}_{17}\text{H}_{12})]^+\), -1.02 eV) before entering the first transition state (TS1, \([\text{C}_{58}(\text{7 C-ring})-(\text{a-C}_{17}\text{H}_{12})]^+\), -0.91 eV) and passing the first activation barrier (0.02 eV). After that, the product (product, \([\text{C}_{58}(\text{7 C-ring})-(\text{a-C}_{17}\text{H}_{12})]^+\), -0.28 eV), which passes the second activation barrier (0.74 eV).

As presented in Figure 5(B), the CH unit (the same side as the CH\(_2\) unit) first bonds to the fullerene cage (7 C-ring); then, the CH unit enters an intermediary state (Inter, \([\text{C}_{58}(\text{7 C-ring})-(\text{b-C}_{17}\text{H}_{12})]^+\), -0.66 eV) before entering the first transition state (TS1, \([\text{C}_{58}(\text{7 C-ring})-(\text{b-C}_{17}\text{H}_{12})]^+\), -0.67 eV) and passing the first activation barrier (0.26 eV). After that, the product (product, \([\text{C}_{58}(\text{7 C-ring})-(\text{b-C}_{17}\text{H}_{12})]^+\), -0.21 eV), which passes the second activation barrier (0.46 eV). Considering the energy loss of the van der Waals molecules, such as IR emissions and collisions, a large number of covalently bonded molecules are formed during the transition states.

During the adduct reaction process, charge transfer (blue line) cannot occur between the \(\text{C}_{58}^+\) cation and 2,3-benzofluorene. As presented in Figure 5, the charge exchange reaction of \(\text{C}_{58}^+\) with 2,3-benzofluorene is endothermic with an energy of 0.32 eV, which is thermodynamically unfavourable.

The charge transfer channel between the \(\text{C}_{58}^+\) cation and 2,3-benzofluorene was also calculated. The charge exchange reaction of \(\text{C}_{58}^+\) with 2,3-benzofluorene is endothermic with an energy of 0.64 eV, which is thermodynamically unfavourable. We note that we simplified the charge transfer channel in the ion-molecular collision reaction process, and our conclusions were mainly based on the DFT calculation results. A molecular dynamic simulation that treats the fullerene cations as metallic spheres and 2,3-benzofluorene as thin circular conducting discs is needed to understand more details of the collision process (Zettergren et al. 2012; Forsberg et al. 2013).

4. Discussion

Overall, the obtained theoretical calculation results are consistent with the experimental results: the charge transfer process is the dominant channel for the reaction system of \(\text{C}_{60}^+\) and 2,3-benzofluorene, while the adduct formation channel has little influence. In contrast, the adduct formation channel is the dominant channel for the reaction system of \(\text{C}_{58}^+\) (7 C-ring) and 2,3-benzofluorene. Thus, the charge transfer process is likely unfavourable for the system of \(\text{C}_{58}^+\) and 2,3-benzofluorene.

In the theoretical calculations, we separate the reaction process between the fullerene cations and 2,3-benzofluorene into two types of channels (Zhen et al. 2019b). The first type of channel is the adduct formation channel, in which two processes occur: the first process is between the fullerene cation and 2,3-benzofluorene to form the van der Waals cluster (Initial), which is energetically unfavourable with the reactants; the second process is the transition from the van der Waals cluster (Initial) to the covalently bonded cluster (Product).

There is a substantially higher energy barrier during the transition from the van der Waals cluster to the covalently bonded cluster for \(\text{C}_{60}^+\) and \(\text{C}_{58}^+\) (6 C-ring) than for \(\text{C}_{58}^+\) (7 C-ring). Thus, if clusters form, fullerene (\(\text{C}_{60}^+\) and \(\text{C}_{58}^+\) (6 C-ring))/2,3-benzofluorene cluster cations are trapped in the van der Waals form and do not have sufficient energy to overcome the energy barrier of the covalently bonded species. In contrast, if they are trapped in van der Waals form, fullerene (\(\text{C}_{58}^+\) (7 C-ring))/2,3-benzofluorene cluster cations can form, overcoming the energy barrier as covalently bonded species. The second type of channel is the charge transfer channel; charge transfer easily occurs between the \(\text{C}_{60}^+\) cation and 2,3-benzofluorene, which is thermodynamically favourable, while charge transfer cannot occur between the \(\text{C}_{58}^+\) cation and 2,3-benzofluorene, which is thermodynamically unfavourable.

According to the theoretical results, we can infer that the formation of van der Waals species is energetically favourable and that all species have relatively strong van der Waals bonding abilities. However, according to the experimental results, almost no clusters based on the fullerene (\(\text{C}_{60}\)) cations form, and we speculate that van der Waals bond cluster molecules do not play an important role in the adduct process. Molecular cluster formation driven by covalent bond forces makes up the overall ion-molecular reaction process. We further surmise that the van der Waals cluster does not survive in our experimental conditions, while the covalently bonded species do.

When reacted with fullerene (\(\text{C}_{60}\) and \(\text{C}_{70}\)) cations, compared with some PAH molecules (e.g. anthracene, \(\text{C}_{14}\text{H}_{10}\)) (Zhen et al. 2019b), 2,3-benzofluorene has a weak covalent bond ability due to the effect of the CH\(_2\) unit. For smaller fullerene molecules, the bonding patterns with anthracene and 2,3-benzofluorene are similar; the cluster is formed mainly due to the higher bonding ability of the smaller fullerene cations.

5. Astronomical implications and conclusions

The total abundance of fullerene (\(\text{C}_{60}\) and \(\text{C}_{70}\)) in interstellar environments, such as the PDRs, is found through several given states as opposed to one (Omont 2016). The abundance of each given state is affected by the surrounding physical and chemical environmental conditions (Berné & Tielens 2012; Omont 2016; Candian et al. 2019). According to previous works, for fullerene species, the cluster formation channels compete with the charge transfer channels. The cluster formation channels are the dominant channels for reactions with singly charged cations, since the charge transfer channel is endothermic for these reactions, while the charge transfer channel are the dominant channels for reactions with doubly and triply charged cations, since the charge transfer channel is exothermic for these reactions (Bohme 2016).

However, based on our current results, charge transfer cations are also the dominant reaction channels for singly charged fullerenes (\(\text{C}_{60/70}\)) when they react with natural 2,3-benzofluorene. Based on this result, for interstellar chemical-evolution models for the molecular spatial evolution of fullerenes and PAH species in space, especially when the concentration ratio of the neutral fullerene molecules to the fullerene monocations is considered, differences in the reaction
of fullerene monocations with their coexisting interstellar neutral molecules should be taken into account, such as anthracene (C\textsubscript{14}H\textsubscript{10})-type molecules, which have higher covalent bond abilities that lead to the formation of clusters (Zhen et al. 2019b), or 2,3-benzofluorene (C\textsubscript{17}H\textsubscript{12})-type molecules, which have weak covalent bond abilities that lead to charge exchanges.

(Bernard-Salas et al. 2012) was unable to satisfactorily explain the relative intensities of the IR emission bands attributed to C\textsubscript{60} in planetary nebulae. They suggested that other substances, such as C\textsubscript{70}, contributed to some of the bands, resulting in the inconsistencies observed in their results. Based on our results, fullerenes (C\textsubscript{60/70}) increase the abundance of their monocation forms through collision reactions with other neutral molecules, which may better explain their evolution and charge distribution characteristics in the ISM. However, further observational and experimental studies are required to address these issues.

In addition, it has been suggested that PAH clusters play an important role in extended red emissions, which are prominent in various interstellar and circumstellar environments (Rhee et al. 2007; Montillaud et al. 2013). The fullerene/PAH adducts formed in our experiments may be relevant to these extended red emissions. Similarly, these types of clusters may play an important role in the IR spectral complexity of circumstellar environments, where C\textsubscript{60} is prominent (Camí et al. 2010; Sloan et al. 2014; Otsuka et al. 2014). Furthermore, the covalent fullerene/PAH clusters obtained in this work may be an important step in the formation of larger carbon grains (Dunk et al. 2013). Our present results indicate that smaller fullerene cations, such as C\textsubscript{38}\textsuperscript{+} and C\textsubscript{56}\textsuperscript{+} (C\textsubscript{68}\textsuperscript{+} and C\textsubscript{66}\textsuperscript{+}), may form adducted clusters with PAH species even though these PAHs have weak covalent bond abilities (e.g. 2,3-benzofluorene, C\textsubscript{17}H\textsubscript{12}). Hence, if these smaller fullerenes are present in space, the formation of covalently bonded fullerene-based clusters could produce an extended family of large molecules (Candian et al. 2018).

In conclusion, the experimental and theoretical results of the gas-phase reaction of fullerene cations with 2,3-benzofluorene are presented, demonstrating the importance of charge transfer channels and illustrating the competition between charge transfer and molecular cluster formation channels. The ion-molecule collision reaction between fullerene cations and 2,3-benzofluorene offers a good model for understanding the physical-chemical processes of the charge transfer channel and cluster formation channels. Thus, fullerenes (C\textsubscript{60/70}) increase the abundance of their monocation forms through collision reactions with other neutral molecules in the ISM. Further observational studies are warranted, especially in consideration of the spatial evolution of fullerenes and PAH species in space.

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Appendix A: Experimental and theoretical calculation methods

The experiment was performed on our apparatus, which was equipped with a quadrupole mass filter, a quadrupole ion trap, and a time-of-flight mass spectrometer, which was described in detail in (Zhen et al. 2019a,b). Briefly, gas-phase fullerene (C\textsubscript{60} and C\textsubscript{70}) was produced by heating fullerene powder to \~ 573 and 593 K and then ionizing the power with an electron gun. These fullerene ions were then transported into the ion trap for accumulation via an ion gate and the quadrupole mass filter. Due to the collision between the buffer gas helium and fullerene cations in the ion trap, we assume that the fullerene cations are all in their ground state and have the same temperature as He (room temperature, 298 K). The gas-phase 2,3-benzofluorene (C\textsubscript{17}H\textsubscript{12}) molecules were produced by heating its powder to \~ 323 K, then continuously effusing its product towards the centre of the trap for a chemical reaction with the fullerene cations. After some amount of time after the reaction, the ions were detected by reflection time-of-flight mass spectrometry. A highly precise digital generator was used to control the time sequences, including the ion gate, the beam shutter, and the measurement of the trapped ions. The experiments were run at 0.1 Hz, with a measurement period lasting 10 s. The following time sequence schemes were employed in the experiments: the ion gate remained open from 0-9.88 s, allowing fullerene cations to accumulate and react with 2,3-benzofluorene. Then, at 9.88 s, all trapped cation species were removed from the ion trap and detected by time-of-flight mass spectrometry.

The calculations were based on density functional theory (DFT) at the B3LYP level (Becke 1992). A basis set of 6-31+g(d,p), including the dispersion correction (D3), was chosen to describe the clusters with van der Waals bonds (Grimme et al. 2011; Sato et al. 2013; Frisch et al. 2016). The geometries of all species were optimized at the local minimum of their potential energy surface in the calculation. For the fullerene cations, 2,3-benzofluorene, and their adducts, frequency calculations were also performed based on their optimized geometries. Furthermore, the frequency calculation determined the zero-point energy and the thermal corrections to the molecular energy. We emphasize that the basis set-superposition error correction is not included in the calculation because of its slight influence on the binding energy.

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